NAME:_____

M&M IONIC COMPOUND LAB

Understanding the ratios of charges to balance compounds can be a difficult task. In an ionic compound, the positive charge must equal the negative charge. You would need two +1 ions to combine with one -2 ion. To better understand the charge ratios we can use models of the compounds. This lab uses M&M's as the models.

Each color of M&M represents a different value of ionic charge:

red	= +1	orange	= -1
blue	= +2	yellow	= -2
brown	= +3	green	= -3

In the boxes below you are to:

1. Write the correct balanced compound for the each of the names provided to you by the instructor .
2. Put the correct numbers and colors of M&M's as a model of the compound to exemplify the charge as zero overall. (Use only one M&M for each polyatomic ion, as well.)

E.g. Sodium oxide would be written Na_2O and would be modeled with two red and one yellow M&M.

Sodium carbonate would be written Na_2CO_3 and would be modeled with two red and one yellow, as well.

1. _____ 2. _____ 3. _____

4. _____ 5. _____ 6. _____

7. _____ 8. _____ 9. _____